**Ions 2**

Today in Science class, Mr. Smith did a demonstration.

Sulfuric Acid

In this demonstration, Mr. Smith filled a small glass beaker with sucrose. He covered his hands in baking powder, or sodium bicarbonate. Then he poured some sulfuric acid into the sucrose. The reason he covered his hands in sodium bicarbonate is because it counteracts the acid because it is a base. When he st*irred the acid/sucrose mixture with a popsicle stick, the stick almost instantaneously burned because the acid was so strong. Then, the sulfuric acid reacted with the sucrose and formed a ton of carbon, so the small amount of sucrose and small amount of acid formed a large pillar of carbon, water, and sulfuric acid.*

*Things I learned:*

* The formula for what happened in the experiment is the following:
  + C12H22O11(s) + H2SO4(aq) → 12C(s) + 11H2O(aq) + H2SO4(aq)
* That formula is an example of chemical change.
* Most atoms that form ions do so when they are dissolved in water.
* When salt is dissolved in water, it seperates into two ions.
* You get free-floating sodium ions and free-floating chlorine ions.
* It is still salt.
* The water will still be salty.
* Acids are things that form hydrogen ions when dissolved in water.
* Bases are things that form hydroxide ions when dissolved in water.
* Strong acids form lots of hydrogen ions.
* Sulfuric acid is car battery acid.
* By state law, teachers and students have to wear chemical goggles around strong acids.
* “Zucker” means sugar in German.

